

4. It has been shown that the assumption about the independence of the partial molar volume from the concentration and the pressure for the system considered does not give rise to errors in excess of the experimental error.

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12. In the general case  $v_2$  denotes the molar volume of the solvent.
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16. If we know the  $\bar{v}_1$  value it is also possible to calculate the  $a^{\frac{1}{2}}/\bar{v}_1$  values from the Van der Waals constants  $a$  when the extrapolation of  $a^{\frac{1}{2}}/v$  becomes unnecessary. (See previous paper.)